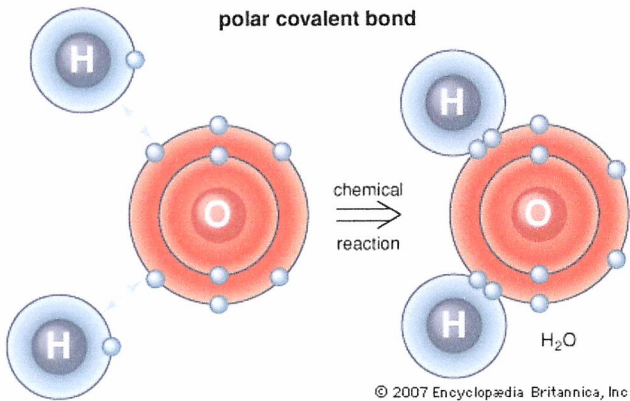


## Consolidation Worksheet on Water

**Water is the most abundant molecule in cells, whole organisms and on earth. This is a consequence of its unique physical and chemical properties.**

### Chemical structure of the molecule



Describe how the hydrogen and oxygen atoms are bonded together

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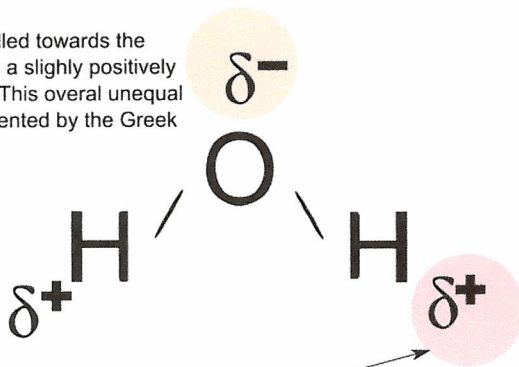
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Are the electrons in the bond shared equally, very unequally or slightly unequally?

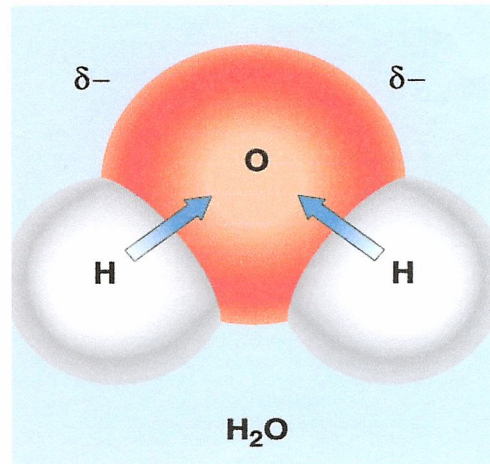
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### Polarity of water molecules

electrons are pulled towards the oxygen, creating a slightly positively charged region. This overall unequal charge is represented by the Greek delta, for dipole



electrons are pulled away from the hydrogen towards the oxygen, creating a slightly positively charged region



Inc., publishing as Benjamin Cummings

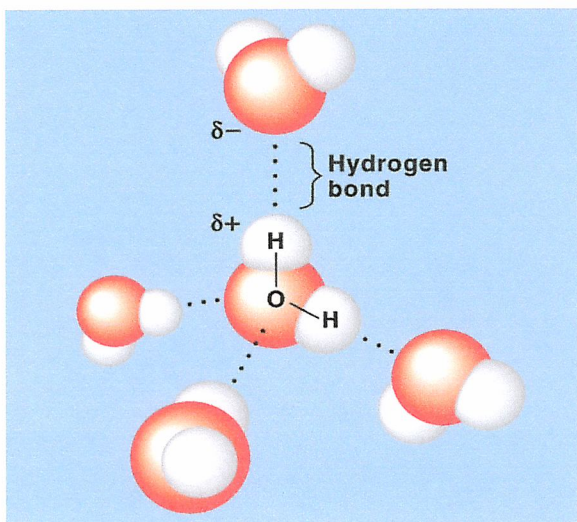
What is the overall charge on a water molecule? -----

Why is the molecule described as a dipole?

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What are H bonds? Explain how Hydrogen bonds form between water molecules.

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How strong are the bonds?

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Draw below one water molecule. Show how it is bonded to three other molecules with intermolecular H bonds. You can be asked to do this in an exam.

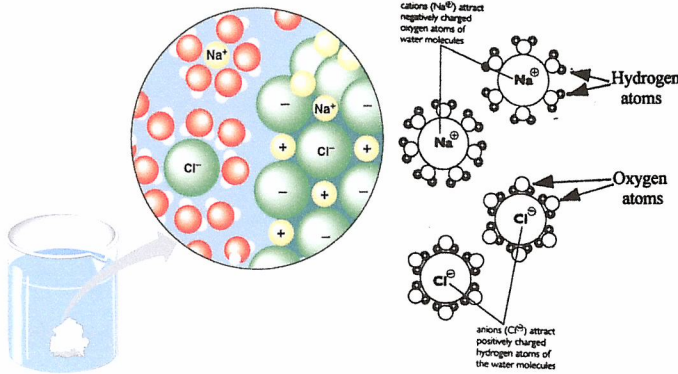
Watch the following video on H bonds <https://www.youtube.com/watch?v=PyC5r2mB4d4> before you do your drawing.

What is the 'sticking' together of water molecules called? -----

Properties that result from the structure

**1. A solvent**

**Water is sometimes described as the universal solvent, allowing chemical reactions to take place in solution.**



What type of substances can water dissolve?

What type of substances can water not dissolve?

In your own words describe how water dissolves an ionic solid such as sodium chloride

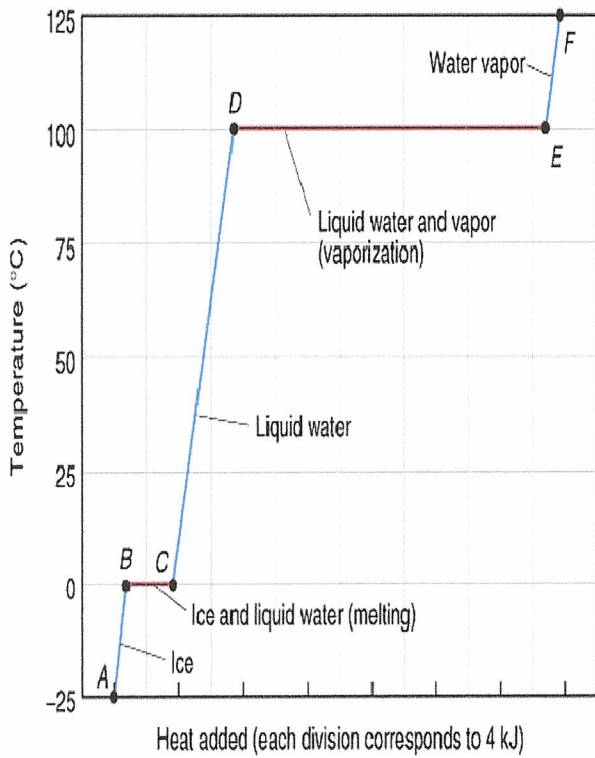
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**2. It has a high latent heat of vaporisation and can act as a coolant**



Give two examples of how water is used as a coolant

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What do we mean by the latent heat of vaporization?

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Why is it so high for water?

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Why is there a cooling effect when surface molecules of water evaporate?

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### 3. Temperature Stabilizer (buffer)

What is specific heat capacity?

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Why does water have such a high specific heat capacity?

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How does this help aquatic environments?

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Why else is this important for organisms?

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#### 4. Thermal Insulator



In aquatic environments why is it important that water is less dense as a solid than it is as a liquid?

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Why is it less dense as a solid?

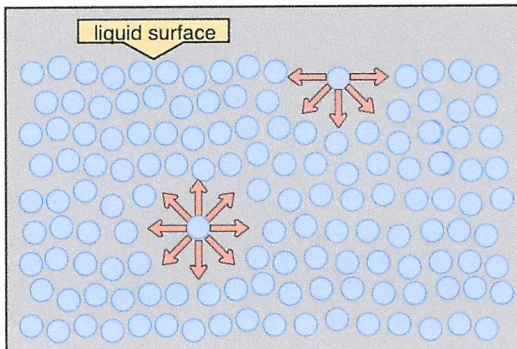
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#### 5. Transport



What property does cohesion lead to that is useful in tall trees?

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What property does cohesion lead to that allows animals such as pond skaters walk on the surface of water

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#### 6. A metabolite

A metabolic reaction is a chemical reaction that happens in a living organism to keep it alive. The metabolism is the sum of all of the chemical reactions together. A metabolite is a substance involved in a metabolic reaction.

How is water involved in photosynthesis? Is it a reactant or a product? Write the word equation below.

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What about respiration?

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In the next topic you will learn about two further metabolic reactions that involve water - **Condensation and Hydrolysis**